

$$\textcircled{12} \text{NH}_4\text{Br} = 97.9426 \text{g}$$

$$\% \text{N} = \frac{14.007 \text{g}}{97.9426 \text{g}} \times 100 = 14.30\% \text{N}$$

$$\% \text{H} = \frac{4.0316 \text{g}}{97.9426 \text{g}} \times 100 = 4.12\% \text{H}$$

$$\% \text{Br} = \frac{79.904 \text{g}}{97.9426 \text{g}} \times 100 = 81.58\% \text{Br}$$

$$\textcircled{13} \text{Ca}(\text{NO}_3)_2 = 164.094 \text{g}$$

$$\% \text{Ca} = \frac{40.08 \text{g}}{164.094 \text{g}} \times 100 = 24.43\% \text{Ca}$$

$$\% \text{N} = \frac{28.014 \text{g}}{164.094 \text{g}} \times 100 = 17.07\% \text{N}$$

$$\% \text{O} = \frac{96 \text{g}}{164.094 \text{g}} \times 100 = 58.50\% \text{O}$$

$$\textcircled{16} \text{Zn}_3(\text{PO}_4)_2 = 386.118 \text{g}$$

$$\% \text{Zn} = \frac{196.17 \text{g}}{386.118 \text{g}} \times 100 = 50.81\% \text{Zn}$$

$$\% \text{P} = \frac{61.948 \text{g}}{386.118 \text{g}} \times 100 = 16.04\% \text{P}$$

$$\% \text{O} = \frac{128 \text{g}}{386.118 \text{g}} \times 100 = 33.15\% \text{O}$$

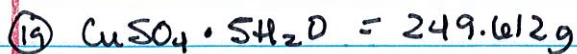
$$\textcircled{18} (\text{NH}_4)_2\text{SO}_4 = 132.1292 \text{g}$$

$$\% \text{N} = \frac{28 \text{g}}{132.1292 \text{g}} \times 100 = 21.19\% \text{N}$$

$$\% \text{H} = \frac{8.0632 \text{g}}{132.1292 \text{g}} \times 100 = 6.10\% \text{H}$$

$$\% \text{S} = \frac{32.066 \text{g}}{132.1292 \text{g}} \times 100 = 24.27\% \text{S}$$

$$\% \text{O} = \frac{64 \text{g}}{132.1292 \text{g}} \times 100 = 48.44\% \text{O}$$



$$\% \text{Cu} = \frac{63.546\text{g}}{249.612\text{g}} \times 100 = 25.46\% \text{Cu}$$

$$\% \text{S} = \frac{32.066\text{g}}{249.612\text{g}} \times 100 = 12.85\% \text{S}$$

$$\% \text{O} = \frac{64\text{g}}{249.612\text{g}} \times 100 = 25.64\% \text{O}$$

$$\% \text{H}_2\text{O} = \frac{90\text{g}}{249.612\text{g}} \times 100 = 36.06\% \text{H}_2\text{O}$$



$$\% \text{P} = \frac{30.974\text{g}}{270.686\text{g}} \times 100 = 11.44\% \text{P}$$

$$\% \text{Br} = \frac{239.712\text{g}}{270.686\text{g}} \times 100 = 88.56\% \text{Br}$$



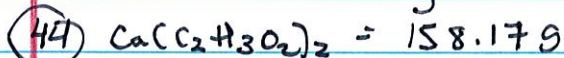
$$\% \text{Sr} = \frac{87.62\text{gSr}}{158.52\text{g}} \times 100 = 55.27\% \text{Sr}$$



$$\% \text{C} = \frac{12.01\text{g}}{100.18\text{g}} \times 100 = 11.99\% \text{C}$$



$$\% \text{H}_2\text{O} = \frac{90\text{g}}{249.612\text{g}} \times 100 = 36.06\% \text{H}_2\text{O}$$



$$\% \text{H} = \frac{6.0474\text{g}}{158.17\text{g}} \times 100 = 3.82\% \text{H}$$



$$\% \text{O} = \frac{64\text{g}}{158.03\text{g}} \times 100 = 40.50\% \text{O}$$